

Cain, Billy

From: Lupher, Chuck
Sent: Tuesday, October 14, 1997 5:22 PM
To: Cain, Billy
Cc: Hermann, Trey; Sommers, Andy
Subject: I'm gone...

Billy,

I need to get some stuff done, and now is a good time. My system is on, so Trey can get to the .raws on my hard drive.

- Talked to Trey, and he's going to have to pull off some Debablizer voodoo to get the frames and clean plate to match. He won't have anything for me to test until later. I voiced my concern with altering the clean plate, possibly causing a mismatch with the other postage stamps. Trey is going to modify a test set first, and I'll look at them. We may need to run ALL of the .raws through the process to make sure that everything matches. Anyway, I'll let Trey do his thing, then look at the results.
- I've let Andy know that I have a few issues that need to be cleared up with him. He will let me know when he's available. The problems I'm getting should (hopefully) be independent from the postage stamp properties info that I have been accumulating. I'll fill you in after I talk to Andy.
- I cut up the B1 group scene, with Hawk a separate postage stamp. It looks a lot better.

Tomorrow, I'm going to make 'still shots' for the problem postage stamps so we'll be ready if they are needed. Also, I'm adding notes to all of the postage stamp properties printouts indicating any unique issues. (Like 0040cz (the new Hawk postage stamp) is meant to be used with 0040cg (the rest of the group)). Hopefully this will help.

See ya tomorrow,

Chuck
<cpl>

1. The first part of the experiment was to determine the effect of temperature on the rate of reaction. The reaction was carried out at three different temperatures: 25°C, 35°C, and 45°C. The rate of reaction was measured by the time taken for the reaction to complete.

2. The second part of the experiment was to determine the effect of concentration on the rate of reaction. The reaction was carried out at three different concentrations: 0.1M, 0.2M, and 0.3M. The rate of reaction was measured by the time taken for the reaction to complete.

3. The third part of the experiment was to determine the effect of surface area on the rate of reaction. The reaction was carried out with three different surface areas: 10 cm², 20 cm², and 30 cm². The rate of reaction was measured by the time taken for the reaction to complete.

